

Worksheet # C59: pH and pOH

1. What does pH mean? _____
2. What does pOH mean? _____
3. $[H^+][OH^-] =$ _____
4. $pH + pOH =$ _____
5. What is the $[H^+]$ of water? _____ What is the pH of water? _____
6. What is the $[OH^-]$ of water? _____ What is the pOH of water? _____
7. If you have a $1.0 \times 10^{-5} M$ HCl solution,
 - a. What is the $[H^+]$? _____
 - b. What is the pH? _____
 - c. What is the pOH? _____
 - d. What is the $[OH^-]$? _____
8. If you have a $.0001 M$ HCl solution,
 - a. What is the $[H^+]$? _____
 - b. What is the pH? _____
 - c. What is the pOH? _____
 - d. What is the $[OH^-]$? _____
9. If you have a $1.0 \times 10^{-3} M$ NaOH solution,
 - a. What is the $[OH^-]$? _____
 - b. What is the pOH? _____
 - c. What is the pH? _____
 - d. What is the $[H^+]$? _____
10. If you have a $.00001 M$ NaOH solution,
 - a. What is the $[OH^-]$? _____
 - b. What is the pOH? _____
 - c. What is the pH? _____
 - d. What is the $[H^+]$? _____